

MT EDUCARE LTD.

ICSE X

SUBJECT : **CHEMISTRY**

PRACTICAL CHEMISTRY

Assignment Sheet

STEP UP ANSWERSHEET

12.

(i) i]

Test	Sodium Chloride Solution	Sodium Nitrate Solution
Conc. H_2SO_4 is added and heated.	Colourless gas evolved which gives dense white fumes with ammonia.	Reddish brown fumes evolved which liberates violet vapours from KI solution.

ii]

Test	Hydrogen Chloride Gas	Hydrogen Sulphide Gas
A rod dipped in ammonium hydroxide is brought near the gas.	Dense white fumes of Ammonium Chloride are formed.	No reaction takes place.

iii]

Test	Calcium Nitrate Solution	Zinc Nitrate Solution
Ammonium Hydroxide is added first dropwise and then in excess.	No precipitation	Gelatinous white ppt of $Zn(OH)_2$ is formed which is soluble in excess of NH_4OH .

iv]

Test	Carbon dioxide Gas	Sulphur dioxide Gas
The gas is passed through acidified solution of orange coloured $K_2Cr_2O_7$ solution.	No. change is observed.	Orange solution of $K_2Cr_2O_7$ turns green.

[2013]

(ii) (A) Carbon monoxide

[2013]

13.

(i)

Sodium Nitrate	Sodium Sulphite
No. effect.	A colourless gas with suffocating odour evolved that turns pink acidified potassium permanganate colourless.

(ii) Starch iodide paper turns blue-black.

(iii) Cation in P is Ca^{2+}

(iv) Gas Q is hydrogen sulphide

[2014]

14. (i) Hydrogen sulphide
(ii) Lead acetate paper turns silvery black.
(iii) (a) A \hat{E} Cl⁻
(b) B \hat{E} CO₃²⁻
(c) C \hat{E} SO₄²⁻
(iv) (a) W \hat{E} Zn²⁺
(b) X \hat{E} Cu²⁺
(c) Y \hat{E} Ca²⁺ or Mg²⁺
(d) Z \hat{E} NH₄⁺

[2015]

15. When barium chloride is added to sodium sulphate, white precipitate is formed.

[2016]

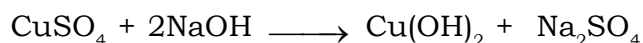
16. (i) B = Non-volatile acid
(ii) A = Typical acid property
(iii) C = Oxidizing agent

[2016]

17. (i) When NH₄OH solution is added drop by drop to copper sulphate solution, a pale blue precipitate is formed which is soluble in excess of NH₄OH and a deep blue or inky blue solution is formed with excess of ammonium hydroxide.



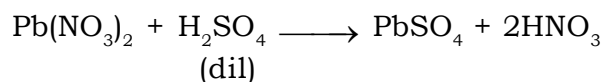
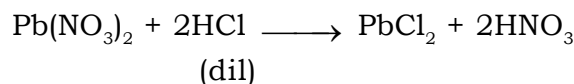
Copper sulphate solution forms a blue precipitate with sodium hydroxide solution. It is insoluble in excess of NaOH.



Copper Sodium Copper Sodium
sulphate hydroxide hydroxide sulphate

- (ii) On adding lead nitrate to both acids we will get a white precipitate. On heating the solution the one whose precipitate will redissolve will be dil.HCl and the one with insoluble precipitate will be dil.H₂SO₄.

Actually on adding lead nitrate to HCl, PbCl₂ precipitates out and on heating the solution it redissolves. But in case of H₂SO₄, PbSO₄ is formed which is insoluble even on heating it.



[2017]

18. (i) Potassium chloride
(ii) Zinc carbonate

[2017]

